they must have a lining of silicious material, owing to the readiness with which carbon is absorbed by the molten material.

The analysis of the product requires a special study of methods. This has been carried out with fairly satisfactory results. The composition of a representative sample was found as shown in the preceding table.

No determination thus far has been made for carbon, boron or titanium.

URBANA, ILLINOIS.

A CORRECTION.

In an article entitled, "Can the Dissociation Theory be Applied to Solid Solutions in Steels?" which appeared in the September number of THIS JOURNAL, the specific resistances of the steel which had been hardened and subsequently reheated to different temperatures were given. In order to render the values as near the absolute value as possible corrections to the observed values necessitated through the calibrations of the recording instruments were made. The corrections expressed in microhms due to the high reading of the ammeter used amount in the present case to from 0.21 to 0.24 microhms, according to the specific resistance. As the ammeter read too high these corrections should have been subtracted from the observed values, but by mistake the corrections were in all cases added. In order, therefore, to reduce the values reported to their true value, from 0.42 to 0.48 microhms should be subtracted from the figures given in the table. This correction would in no way affect the significance of the results or any deductions drawn therefrom.

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HEXABROMODIACETYL.

BY C. LORING JACKSON AND ROGER ADAMS.

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In an earlier paper[1] A. H. Fiske and one of us described two acids, one melting at 207°, the other at 174°, made by the action of sodium hydroxide on tetrabromo-*o*-quinone, and it was also stated that, when treated with bromine and water, these acids were converted into a yellow diketone, to which the formula $CBr_3CHBrCOCOCBr_3$ was assigned. On continuing the study of this compound we have found that this formula is not the true one, but that it is really the hexabromodiacetyl $CBr_3COCOCBr_3$. The data on which this conclusion rests are given in Table I containing all the analytical results obtained by us from the substance itself and its principal derivatives. The first column gives the percentages calculated for $CBr_3CHBrCOCOCBr_3$, the second the data obtained by

[1] *Am. Chem. J.*, 50, 341 (1913).